

Half-life:

The half-life of a radioactive element (substance) is defined as the time interval during which the number of atoms decreases to half of the initial number.

If $T_{1/2}$ is the half-life period of a radioactive substance then

$$\text{For, } t = T_{1/2}, \quad N = \frac{N_0}{2}$$

$$\text{Since, } N = N_0 e^{-\lambda t}$$

$$\therefore \frac{N_0}{2} = N_0 e^{-\lambda T_{1/2}}$$

$$\text{or, } \frac{1}{2} = e^{-\lambda T_{1/2}}$$

$$\text{or, } 2 = e^{\lambda T_{1/2}}$$

Taking \ln on both sides, we get

$$\ln 2 = \lambda T_{1/2}$$

$$\text{or, } T_{1/2} = \frac{\ln 2}{\lambda}$$

$$\therefore T_{1/2} = \frac{0.693}{\lambda}$$

: Relation between half-life and decay constant.

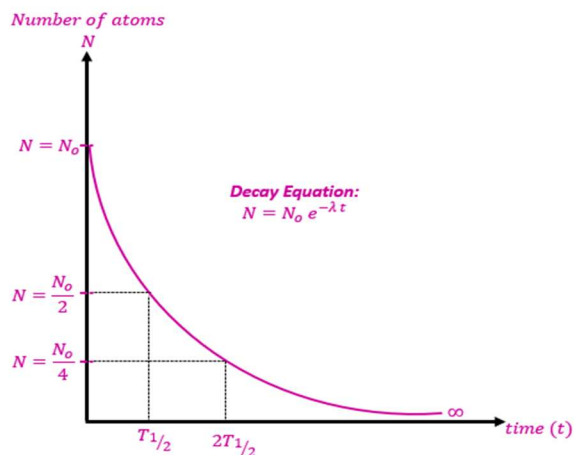


Figure: Decay Curve

Mean-Life (\bar{T}):

The mean life of a radioactive substance is defined as the ratio of sum of lives of all radioactive atoms to total number of radioactive atoms.

$$\text{i.e., mean life} = \frac{\text{sum of lives of all radioactive atoms}}{\text{total number of radioactive atoms}}$$

Mathematically,

$$\checkmark \text{ mean life} = \frac{1}{\lambda}$$

$$\checkmark \text{ mean life} = \sqrt{2} T_{1/2}$$

Radioactive decay constant: [Definition of λ]

Since,

$$\frac{dN}{dt} = \lambda N \quad (\text{magnitude only})$$

$$\therefore \lambda = \frac{dN/dt}{N}$$

Thus, decay constant of a radioactive substance is defined as the ratio of the rate of disintegration of radioactive atoms at a given instant of time to the number of atoms present at that instant.

$$\text{Also, } N = N_0 e^{-\lambda t}$$

$$\text{If, } \lambda = \frac{1}{t}, \text{ (reciprocal of time)}$$

$$\text{Then, } N = N_0 e^{-\frac{1}{t} \times t}$$

$$N = N_0 e^{-1}$$

$$N = 0.368 N_0$$

$$N = 37\% \text{ of } N_0$$

OR,

Number of radioactive substances decreases by about 63% of their original number.

Hence the radioactive decay constant may also be defined as the reciprocal of time in which the number of atoms of a radioactive substance decreases to about 37% of their original (initial) number.